Acta Crystallographica Section E
Structure Reports Online
ISSN 1600-5368

2-Benzyliminomethyl-4-chlorophenol

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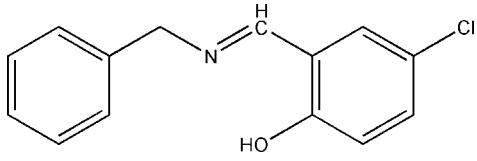
Received 17 December 2007; accepted 4 February 2008

Key indicators: single-crystal X-ray study; T = 298 K; mean σ(C–C) = 0.006 Å; R factor = 0.045; wR factor = 0.116; data-to-parameter ratio = 14.1.

The title compound, $C_{14}H_{12}ClNO$, is a Schiff base derived from the condensation of equimolar quantities of 5-chlorosalicylaldehyde and 1-benzylamine. The molecule has a *trans* configuration with respect to the imine C=N double bond. The N atom is involved in an intramolecular O—H···N hydrogen bond.

Related literature

For related literature, see: Ali *et al.* (2002); Cukurovali *et al.* (2002); Tarafder *et al.* (2002).



Experimental

Crystal data

$C_{14}H_{12}ClNO$
M_r = 245.70
Monoclinic, $P2_1/c$
a = 14.3693 (18) Å
b = 6.0401 (8) Å
c = 14.777 (2) Å
β = 103.911 (2)°
V = 1244.9 (3) Å^3
Z = 4
Mo $K\alpha$ radiation
μ = 0.29 mm^{-1}
T = 298 (2) K
0.52 × 0.38 × 0.11 mm

Data collection

Bruker SMART APEX diffractometer
Absorption correction: multi-scan (*SADABS*; Sheldrick, 1996)
T_{min} = 0.864, T_{max} = 0.969
5203 measured reflections
2177 independent reflections
864 reflections with $I > 2\sigma(I)$
R_{int} = 0.046

Refinement

$R[F^2 > 2\sigma(F^2)]$ = 0.045
$wR(F^2)$ = 0.116
S = 0.98
2177 reflections
154 parameters
H-atom parameters constrained
$\Delta\rho_{max}$ = 0.15 e Å^{-3}
$\Delta\rho_{min}$ = −0.22 e Å^{-3}

Table 1
Hydrogen-bond geometry (Å, °).

D—H···A	D—H	H···A	D···A	D—H···A
O1—H1···N1	0.82	1.87	2.597 (4)	148

Data collection: *SMART* (Bruker, 2000); cell refinement: *SMART*; data reduction: *SAINT* (Bruker, 2000); program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *SHELXTL* (Sheldrick, 2008); software used to prepare material for publication: *SHELXTL*.

The authors thank Baoji University of Arts and Sciences for support.

Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: CS2068).

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supplementary materials

Acta Cryst. (2008). E**64**, o726 [**doi:10.1107/S1600536808003802**]

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X. Zhang and Z. Li

Comment

Schiff base compounds have been of great interest for many years. These compounds played important role in the development of coordination chemistry related to catalysis and enzymatic reactions, magnetism and molecular architectures. These properties stimulated our interest in this field. The title compound was obtained as a new antipyrine Schiff base.

Its molecular structure and a crystal packing are illustrated in Figs.1 and 2, respectively. Atom N1 is a bridging N atom linking the two parts of the compound. The dihedral angle between the two phenyl rings is 72.91 (9) °. In the crystal structure, there exists an intramolecular O—H—N hydrogen bond involving hydroxyl atom O1 and imine atom N1 (Table 1).

Experimental

All reagents used were of analytical grade from commercial sources and used without further purification. 5-Chlorosalicylaldehyde (0.1 mmol, 15.65 mg) and 1-benzylamine (0.1 mmol, 10.7 mg) were dissolved in methanol (10 ml). The resulting solution was stirred for 30 min, filtered and the filtrate allowed to stand at room temperature. Yellow crystals of the title compound appeared after two weeks of slow evaporation of the solvent.

Refinement

All H atoms were placed in geometrically idealized positions and constrained to ride on their parent atoms with C—H distances in the range 0.93–0.96 Å and U_{iso}(H) = 1.2U_{eq} or 1.5U_{eq}(C/O)

Figures

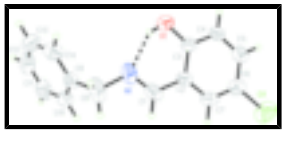

Fig. 1. The structure of the title compound with 30% probability displacement ellipsoids. H atoms are shown as spheres of arbitrary radii. The dotted line represent a hydrogen bond.

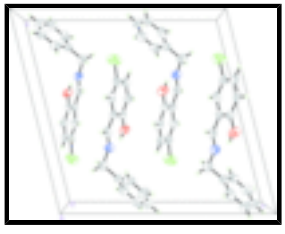

Fig. 2. Packing of the molecules viewed along the *b* axis.

2-Benzyliminomethyl-4-chlorophenol

Crystal data

$C_{14}H_{12}ClNO$	F_{000} = 512
M_r = 245.70	D_x = 1.311 Mg m^{-3}
Monoclinic, $P2_1/c$	Mo $K\alpha$ radiation λ = 0.71073 Å
a = 14.3693 (18) Å	Cell parameters from 877 reflections
b = 6.0401 (8) Å	θ = 2.8–25.1º
c = 14.777 (2) Å	μ = 0.29 mm^{-1}
β = 103.911 (2)º	T = 298 (2) K
V = 1244.9 (3) Å^3	Rod, yellow
Z = 4	0.52 × 0.38 × 0.11 mm

Data collection

Bruker SMART APEX diffractometer	2177 independent reflections
Radiation source: fine-focus sealed tube	864 reflections with $I > 2\sigma(I)$
Monochromator: graphite	R_{int} = 0.046
T = 298(2) K	θ_{max} = 25.0º
φ and ω scans	θ_{min} = 1.5º
Absorption correction: multi-scan (SADABS; Sheldrick, 1996)	h = −17→17
T_{min} = 0.864, T_{max} = 0.969	k = −7→6
5203 measured reflections	l = −17→10

Refinement

Refinement on F^2	Secondary atom site location: difference Fourier map
Least-squares matrix: full	Hydrogen site location: inferred from neighbouring sites
$R[F^2 > 2\sigma(F^2)]$ = 0.045	H-atom parameters constrained
$wR(F^2)$ = 0.116	$w = 1/[\sigma^2(F_o^2) + (0.0248P)^2 + 0.4795P]$ where $P = (F_o^2 + 2F_c^2)/3$
S = 0.98	$(\Delta/\sigma)_{max} < 0.001$
2177 reflections	$\Delta\rho_{max}$ = 0.15 e Å^{-3}
154 parameters	$\Delta\rho_{min}$ = −0.22 e Å^{-3}
Primary atom site location: structure-invariant direct methods	Extinction correction: none

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (Å^2)

	x	y	z	U_{iso}*/U_{eq}
Cl1	0.24166 (8)	0.1883 (2)	0.07747 (9)	0.1401 (6)

N1	0.6895 (2)	0.0444 (5)	0.2091 (2)	0.0878 (10)
O1	0.60143 (16)	−0.2928 (4)	0.11796 (16)	0.0884 (8)
H1	0.6468	−0.2203	0.1478	0.133*
C1	0.6093 (3)	0.1388 (6)	0.1982 (2)	0.0772 (11)
H1A	0.6070	0.2803	0.2223	0.093*
C2	0.5198 (2)	0.0337 (6)	0.1488 (2)	0.0580 (8)
C3	0.5196 (3)	−0.1769 (6)	0.1098 (2)	0.0613 (9)
C4	0.4338 (3)	−0.2717 (6)	0.0612 (2)	0.0738 (10)
H4	0.4337	−0.4118	0.0350	0.089*
C5	0.3499 (3)	−0.1585 (7)	0.0520 (2)	0.0769 (11)
H5	0.2927	−0.2222	0.0194	0.092*
C6	0.3489 (3)	0.0483 (7)	0.0904 (2)	0.0740 (10)
C7	0.4331 (3)	0.1431 (6)	0.1383 (2)	0.0717 (10)
H7	0.4320	0.2831	0.1642	0.086*
C8	0.7753 (3)	0.1670 (8)	0.2583 (3)	0.1196 (16)
H8A	0.8064	0.0884	0.3147	0.143*
H8B	0.7567	0.3121	0.2759	0.143*
C9	0.8432 (2)	0.1917 (8)	0.1974 (3)	0.0717 (10)
C10	0.8429 (3)	0.3790 (7)	0.1443 (3)	0.0875 (12)
H10	0.8002	0.4932	0.1469	0.105*
C11	0.9053 (3)	0.3980 (8)	0.0877 (3)	0.1031 (15)
H11	0.9042	0.5248	0.0517	0.124*
C12	0.9683 (3)	0.2347 (11)	0.0834 (3)	0.1099 (16)
H12	1.0110	0.2503	0.0454	0.132*
C13	0.9693 (3)	0.0497 (9)	0.1342 (4)	0.1057 (15)
H13	1.0120	−0.0640	0.1308	0.127*
C14	0.9075 (3)	0.0294 (7)	0.1905 (3)	0.0904 (12)
H14	0.9091	−0.0991	0.2255	0.108*

Atomic displacement parameters (Å^2)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
Cl1	0.1065 (8)	0.1452 (11)	0.1546 (11)	0.0485 (8)	0.0038 (7)	0.0075 (9)
N1	0.0742 (19)	0.117 (3)	0.075 (2)	−0.030 (2)	0.0236 (18)	−0.004 (2)
O1	0.0883 (17)	0.0730 (17)	0.113 (2)	0.0039 (14)	0.0419 (15)	−0.0134 (15)
C1	0.102 (3)	0.080 (3)	0.056 (2)	−0.032 (3)	0.030 (2)	−0.014 (2)
C2	0.078 (2)	0.051 (2)	0.050 (2)	−0.007 (2)	0.0238 (18)	0.0015 (18)
C3	0.078 (2)	0.052 (2)	0.062 (2)	−0.002 (2)	0.033 (2)	0.0004 (19)
C4	0.096 (3)	0.057 (2)	0.078 (3)	−0.015 (2)	0.038 (2)	−0.015 (2)
C5	0.084 (3)	0.091 (3)	0.055 (2)	−0.016 (2)	0.016 (2)	−0.007 (2)
C6	0.082 (3)	0.077 (3)	0.063 (3)	0.014 (2)	0.017 (2)	0.011 (2)
C7	0.102 (3)	0.053 (2)	0.061 (2)	0.001 (2)	0.023 (2)	−0.0014 (19)
C8	0.092 (3)	0.186 (5)	0.086 (3)	−0.059 (3)	0.031 (3)	−0.032 (3)
C9	0.065 (2)	0.079 (3)	0.068 (3)	−0.016 (2)	0.011 (2)	−0.012 (2)
C10	0.072 (3)	0.073 (3)	0.111 (4)	0.005 (2)	0.011 (2)	−0.004 (3)
C11	0.094 (3)	0.099 (4)	0.111 (4)	−0.022 (3)	0.015 (3)	0.030 (3)
C12	0.079 (3)	0.160 (5)	0.096 (4)	−0.018 (3)	0.029 (3)	−0.002 (4)
C13	0.096 (3)	0.111 (4)	0.105 (4)	0.024 (3)	0.013 (3)	−0.021 (3)

C14	0.114 (3)	0.074 (3)	0.075 (3)	−0.004 (3)	0.007 (3)	0.003 (2)

Geometric parameters (Å, °)

Cl1—C6	1.727 (3)	C7—H7	0.9300
N1—C1	1.260 (4)	C8—C9	1.485 (4)
N1—C8	1.471 (4)	C8—H8A	0.9700
O1—C3	1.349 (3)	C8—H8B	0.9700
O1—H1	0.8200	C9—C14	1.368 (5)
C1—C2	1.462 (4)	C9—C10	1.376 (5)
C1—H1A	0.9300	C10—C11	1.370 (5)
C2—C7	1.385 (4)	C10—H10	0.9300
C2—C3	1.396 (4)	C11—C12	1.351 (5)
C3—C4	1.392 (4)	C11—H11	0.9300
C4—C5	1.365 (4)	C12—C13	1.344 (5)
C4—H4	0.9300	C12—H12	0.9300
C5—C6	1.373 (4)	C13—C14	1.361 (5)
C5—H5	0.9300	C13—H13	0.9300
C6—C7	1.372 (4)	C14—H14	0.9300
C1—N1—C8	117.9 (4)	N1—C8—H8A	109.6
C3—O1—H1	109.5	C9—C8—H8A	109.6
N1—C1—C2	122.4 (4)	N1—C8—H8B	109.6
N1—C1—H1A	118.8	C9—C8—H8B	109.6
C2—C1—H1A	118.8	H8A—C8—H8B	108.1
C7—C2—C3	118.4 (3)	C14—C9—C10	117.2 (4)
C7—C2—C1	120.6 (3)	C14—C9—C8	121.8 (5)
C3—C2—C1	121.0 (3)	C10—C9—C8	121.0 (4)
O1—C3—C4	118.5 (3)	C11—C10—C9	120.3 (4)
O1—C3—C2	121.4 (3)	C11—C10—H10	119.9
C4—C3—C2	120.1 (3)	C9—C10—H10	119.9
C5—C4—C3	119.8 (3)	C12—C11—C10	120.8 (4)
C5—C4—H4	120.1	C12—C11—H11	119.6
C3—C4—H4	120.1	C10—C11—H11	119.6
C4—C5—C6	120.8 (3)	C13—C12—C11	119.9 (5)
C4—C5—H5	119.6	C13—C12—H12	120.1
C6—C5—H5	119.6	C11—C12—H12	120.1
C7—C6—C5	119.8 (3)	C12—C13—C14	119.7 (5)
C7—C6—Cl1	120.4 (3)	C12—C13—H13	120.2
C5—C6—Cl1	119.9 (3)	C14—C13—H13	120.2
C6—C7—C2	121.1 (3)	C13—C14—C9	122.2 (4)
C6—C7—H7	119.4	C13—C14—H14	118.9
C2—C7—H7	119.4	C9—C14—H14	118.9
N1—C8—C9	110.2 (3)		
C8—N1—C1—C2	178.8 (3)	C3—C2—C7—C6	−0.4 (5)
N1—C1—C2—C7	180.0 (3)	C1—C2—C7—C6	178.5 (3)
N1—C1—C2—C3	−1.2 (5)	C1—N1—C8—C9	−121.1 (4)
C7—C2—C3—O1	−179.6 (3)	N1—C8—C9—C14	−84.3 (4)
C1—C2—C3—O1	1.5 (5)	N1—C8—C9—C10	94.8 (4)
C7—C2—C3—C4	0.4 (4)	C14—C9—C10—C11	−0.1 (5)

C1—C2—C3—C4	−178.4 (3)	C8—C9—C10—C11	−179.2 (3)
O1—C3—C4—C5	179.8 (3)	C9—C10—C11—C12	−0.5 (6)
C2—C3—C4—C5	−0.2 (5)	C10—C11—C12—C13	0.9 (7)
C3—C4—C5—C6	−0.1 (5)	C11—C12—C13—C14	−0.9 (7)
C4—C5—C6—C7	0.2 (5)	C12—C13—C14—C9	0.3 (6)
C4—C5—C6—Cl1	179.8 (3)	C10—C9—C14—C13	0.2 (5)
C5—C6—C7—C2	0.1 (5)	C8—C9—C14—C13	179.3 (3)
Cl1—C6—C7—C2	−179.6 (2)		

Hydrogen-bond geometry (Å, °)

D—H···*A*	*D*—H	H···*A*	*D*···*A*	*D*—H···*A*
O1—H1···N1	0.82	1.87	2.597 (4)	148

Fig. 1

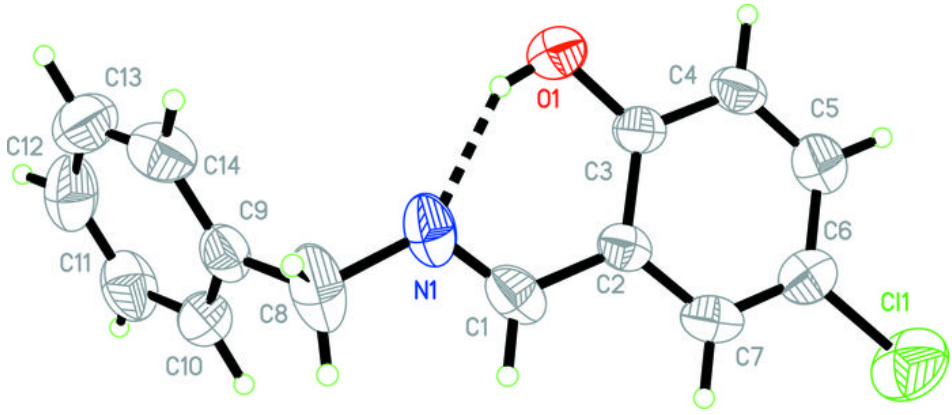

Fig. 2